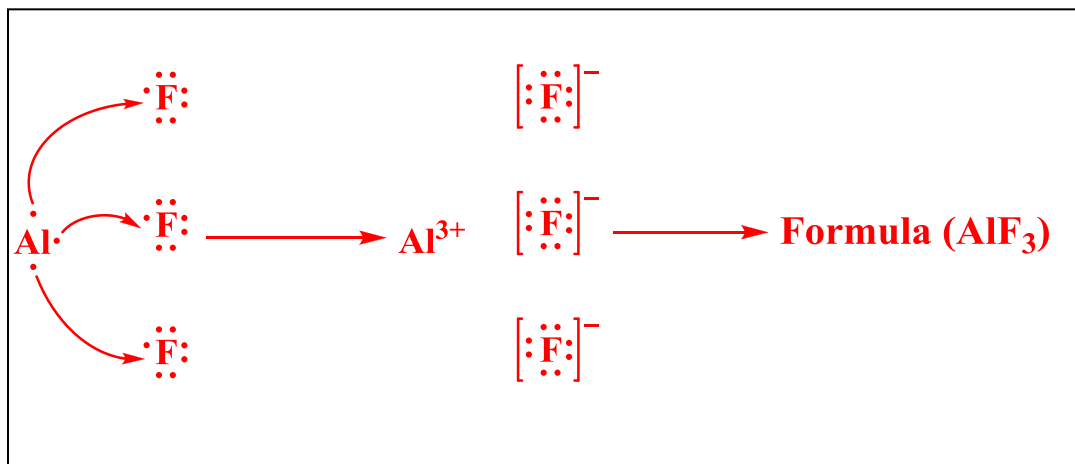


### Practice 3-1

Use the electron-dot symbols to write the equation for the formation of the ionic compound formed between aluminum and fluorine.

Answer



### Practice 3-2

Name the following binary ionic compounds:

BaO    Ca<sub>3</sub>P<sub>2</sub>    Sr<sub>3</sub>N<sub>2</sub>    Ag<sub>2</sub>S    LiBr    NiCl<sub>2</sub>

Answer

BaO	<b>barium oxide</b>	Ag <sub>2</sub> S	<b>silver sulfide</b>
Ca <sub>3</sub> P <sub>2</sub>	<b>calcium phosphide</b>	LiBr	<b>lithium bromide</b>
Sr <sub>3</sub> N <sub>2</sub>	<b>strontium nitride</b>	NiCl <sub>2</sub>	<b>nickel chloride</b>

**Practice 3-3**

Name the following binary ionic compounds:

**Answer**

SnS <sub>2</sub>	<b>tin(IV) sulfide</b>	CuCl <sub>2</sub>	<b>copper(II) chloride</b>
PbI <sub>2</sub>	<b>lead(II) iodide</b>	FeN	<b>iron(III) nitride</b>
Hg <sub>2</sub> O	<b>mercury(I) oxide</b>	Co <sub>2</sub> O <sub>3</sub>	<b>cobalt(III) oxide</b>

**Practice 3-4**

Name the following polyatomic ionic compounds:

**Answer**

Sr(ClO <sub>4</sub> ) <sub>2</sub>	<b>strontium perchlorate</b>
Na <sub>2</sub> CrO <sub>4</sub>	<b>sodium chromate</b>
KH <sub>2</sub> PO <sub>4</sub>	<b>potassium hydrogen phosphate</b>
Ag <sub>2</sub> SO <sub>4</sub>	<b>silver sulfate</b>
Co(OH) <sub>3</sub>	<b>cobalt(III) hydroxide</b>
Cu(C <sub>2</sub> H <sub>3</sub> O <sub>2</sub> ) <sub>2</sub>	<b>copper(II) acetate</b>

**Practice 3-5**

Name each of the following

**Answer**

$S_2F_{10}$	<b>disulfur decafluoride</b>
$SiI_4$	<b>silicon tetraiodide</b>
$P_2O_5$	<b>diphosphorous pentoxide</b>
$B_4Cl_4$	<b>tetraboron tetrachloride</b>
$P_4S_7$	<b>tetraphosphorous heptasulfide</b>
$NBr_3$	<b>nitrogen tribromide</b>
$I_2Cl_6$	<b>diiodine hexachloride</b>
$SI_5$	<b>sulfur pentaiodide</b>