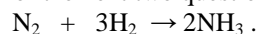


1. The formula mass of C_6H_6 is
A. 78.04 amu B. 78.08 amu C. 78.12 amu D. 78.06 amu
2. What is the number of mol in 6.41 g of SO_2 ?
A. 0.100 mol B. 0.133 mol C. 0.266 mol D. 0.050 mol
3. What is the mass of 2.10 mol of helium?
A. 4.10 g B. 8.40 g C. 4.20 g D. 2.00 g
4. What coefficient is placed in front of H_2O to complete the balancing of the following equation?
 $2Al(OH)_3 + 3H_2SO_4 \rightarrow Al_2(SO_4)_3 + ?H_2O$
A. 6 B. 3 C. 8 D. 2
5. The reaction ($4HNO_3 \rightarrow 4NO_2 + 2H_2O + O_2$), is an example of a _____ reaction.
A. combination B. double replacement C. decomposition D. single replacement
6. In the reaction ($Sn + CuCl_2 \rightarrow SnCl_2 + Cu$), what is reduced?
A. Sn B. Cl^- C. Cu D. Cu^{2+}
7. The reaction ($A + B \rightarrow C + 75 \text{ kcal}$) is an example of a(n) _____ reaction.
A. endothermic B. exothermic C. equilibrium D. slow
8. At equilibrium the rate of the forward reaction is _____ the rate of _____ reaction.
A. the same as, reverse B. smaller than, reverse C. larger than, reverse D. the same as, activation

For the next two questions (9 and 10), consider the following balanced equation:



9. When 4.50 mole of H_2 react, how many grams of N_2 are needed?
A. 28.0 g B. 14.0 g C. 28.5 g D. 42.0 g
10. How many grams of H_2 are required to produce 27.0 g of NH_3 ?
A. 3.70 g B. 6.80 g C. 4.80 g D. 3.20 g

Answers

1	2	3	4	5	6	7	8	9	10
C	A	B	A	C	D	B	A	D	C